Acetone

The result of interview with Mr. Lalu Shafwan Hadi El Wathan. The topic is Acetone.

**What is the General information about acetone?**

Acetone is one of the strongest solvents, because acetone is semi polar. Usually people know that acetone is used to remove nail polish on the nail. Acetone has a low boiling point of 56.6 degrees Celsius. The meaning is at room temperature shouldn’t be left open because acetone will evaporate faster.

Acetone has a ketone functional group and methyl on its right and left side. The functional group located on carbon number 2 from left or right. This causes the compound to be semi-polar. Why is it semi-polar? Because acetone can react with water and non-polar compound which has a short carbon chain.

**In your opinion what is the main problem about the uses of acetone?**

The main problem of using acetone is volatile because acetone has the low boiling point, therefore acetone will be volatile and easy to breathe. Acetone has ketone functional group, there is carbonyl and attach 2 methyl in the right and left side .Therefore, this compound is semi polar. Acetone can react with water and form hydrogen bonds with water and form a bond of Vander walls with non-polar compounds that has short chain. Acetone has a lot of applications, one of them in the chemistry, it can be use as surfactant (as a bridge) but can’t be said surfactant because it can’t dissolve on the long chain. So acetone can be uses as bridge if that compound has short chain.

For the safety, when you want to work with acetone or another compound that has low boiling point, don’t forget to use mask and respirator, because acetone is volatile.

**What is your suggestion about the uses of acetone?**

The suggestion form the Mr. Wathan is if it still fits the function and working on GLP/K3 it is enough. Acetone has low burning point about -17 degrees Celsius, and acetone is flammable, and must be closed due to volatily.